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Some examples of bases are: (i) Sodium hydroxide (NaOH) or caustic soda used in washing soaps. (ii) Potassium hydroxide (KOH) or potash used in bathing soaps. (iii) Calcium hydroxide (Ca(OH)<sub>2</sub>) or lime water used in white wash. (iv) Magnesium hydroxide (Mg(OH)<sub>2</sub>) or milk of magnesia used to control acidity.

## **Notes ACIDS, BASES AND SALTS**

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titrations) to determine their relative pH.

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Bronsted-Lowry base = proton acceptor ( $H^+$  = proton) Bronsted-Lowry acid-base reaction = reaction involving the transfer of a proton e.g.  $HCl + NaOH \rightarrow NaCl + H_2O$   $H^+$  donates  $H^+$  ? acid  $OH^-$  accepts  $H^+$  ? base e.g.  $HNO_3 + NH_3 \rightarrow NH_4^+ + NO_3^-$   $H^+$  donates  $H^+$  ? acid  $NH_3$  accepts  $H^+$  ? base

TASK 1 – Bronsted-Lowry acids & bases

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Acid and Base Worksheet - Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a)  $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$ .  $\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other. b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$

## **Acid and Base Worksheet - Answers**

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Preparing soluble salts, Preparing insoluble salts, Strong and weak acids and alkalis. The following table shows the neutralisation equations when acid reacts with metal, base, metal oxide, metal carbonate, ammonia, and when base reacts with ammonium salt, non-metal oxide.

**Acid, Bases, Salts - IGCSE Chemistry (solutions, examples ...**  
Worksheet: Acids, Bases, and Salts Name\_\_\_\_\_ What are Acids, Bases, and Salts? The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases. ? An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form  $H^+$  ions.

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3) Definitions of acids, bases and salts (i) Acids Eg  $\text{HCl(g)}$  ?

$\text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq})$  A mixture of  $\text{HCl}$  and water contains  $\text{H}^+$  ions and is known as hydrochloric acid Eg  $\text{H}_2\text{SO}_4(\text{l})$  ?  $2\text{H}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$

A mixture of  $\text{H}_2\text{SO}_4$  and water contains  $\text{H}^+$  ions and is

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known as sulfuric acid Eg  $\text{HNO}_3(l) \rightarrow \text{H}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$  A mixture of  $\text{HNO}_3$

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6 ranging from 2002 to the present. Contents include: Acids, Bases and Salts Acids and alkali...

## **Unit 7: Acids, Bases & Salts | Teaching Resources**

1. acids 2. citric acid 3. Base 4. calcium sulphate hemihydrate,  $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$  5. yellow colour, pink colour 6. Hydrochloric Acid (HCl) 7. Acetic Acid ( $\text{CH}_3\text{COOH}$ ) 8. Sodium Hydroxide (NaOH) 9. Ammonium Hydroxide ( $\text{NH}_4\text{OH}$ ) 10. sodium carbonate,  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$ , permanent hardness 11. PH 12. Acid 13. Base 14. Neutral 15. decreases 16. Salt and Water 17. Calcium Carbonate 18. 5.5

## **CBSE Class 10 - Science - CH2 - Acids, Bases and Salts ...**

Chapter 19 1 Acids Bases And Salts Worksheet Answers Acids,

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bases and salts. 7th Grade Science Worksheets, Answer key and Study Guides. Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called carbonates.

## **Acids Bases Salts Answer**

SC.8.P.8.8: Identify basic examples of and compare and classify the properties of compounds, including acids, bases, and salts. Belongs to: Properties of Matter. Join Chef Ragu as he learns about acids, bases, and salts while he cooks up something special. In this interactive tutorial you will compare and classify the properties of compounds that are acids, bases, and salts and identify basic examples of these compounds.

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