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chemical reaction engineering is concerned with the exploitation of
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This reaction follows the equation: $N_2 + 3H_2 \rightarrow 2NH_3$
 $H_0 = 92 \text{ kJ mol}^{-1}$
 $S_0 = 192 \text{ J mol}^{-1} \text{ K}^{-1}$
To find the Gibbs free energy of formation at room temperature, recall that $G_0 = H_0 - TS_0$
 $G_0 = 92 \text{ kJ mol}^{-1} - (298 \text{ K})(0.192 \text{ kJ mol}^{-1} \text{ K}^{-1}) = 35 \text{ kJ mol}^{-1}$
Alternatively, one can find the temperature for which $G = 0$, $T = H_0 / S_0$.

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1 Basic Problems of Chemical Reaction Engineering and ...

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Levenspiel, O (1972) Chemical Reaction Engineering, John ...
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