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Review Sheet

Unit 7 Answer

Key

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AP Chemistry Unit 7  
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~~Chemistry Segment 3~~

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Chem 110 Unit 7

Segment 7AP

**Chemistry Unit 7**

**FRQs Review** *Unit 7*

*Endothermic*

*reaction/graph*

**Chemistry, class-12,**

**unit-7 Chemistry**

~~Review Sheet Unit 7~~

are representative of

each type of problem

worked in this unit. For

more practice on a given



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type of problem, refer back to Note Taking Guides, worksheets, and quizzes. 1. Calculate the molar mass for ethane,  $C_2H_6$ .  $C_2H_6$   $2(12.0 \text{ g}) + 6(1.0 \text{ g}) = 30.0 \text{ g}$  2. Convert 45 g Se to moles of Se.  $? \text{ mol Se} = 45 \text{ g Se} \div 1 \text{ mol Se } 79.0 \text{ g Se} = 0.57 \text{ mol Se}$  3.

~~Review Sheet: Unit 7~~

~~Name Vocabulary: Fill~~

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## Chemistry

~~in the blanks ...~~

are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes.

1. Calculate the molar mass for ethane,  $C_2H_6$ .
2. Convert 45 g Se to moles of Se.
3. Convert  $4.77 \times 10^{24}$  molecules of  $SO_2$  to

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UNIT 7 REVIEW

SHEET –

EQUILIBRIUM •

EQUILIBRIUM is the point at which the rate of the forward reaction is equal to the rate of the reverse reaction in a reversible process. o

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## Chemistry

Reaction is dynamic – still occurring – even though to the eye, it looks like nothing is changing at equilibrium.

o

Observable/measurable quantities

(concentration, pressure, color) are constant at

equilibrium. o The

amounts of products and reactants are typically not equal at equilibrium!

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~~UNIT 7 REVIEW~~

~~SHEET.pdf - UNIT 7~~

~~REVIEW SHEET~~

~~\u20132013 ...~~

CHEMISTRY CP

Name: \_\_\_\_\_ Per:

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Unit 7 Review Sheet:

Thermochemistry KEY

Potential Energy

Diagrams: Choose the appropriate letter. 1.

What is the  $\Delta H$  of the

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forward reaction? f 2.  
What is the activation  
energy of the forward  
reaction? b 3. What is  
the H of the reverse  
reaction? f 4.

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During the first year of high school chemistry or the first semester of college chemistry, all the terms, units of measurement, and atoms, molecules, elements, and compounds may seem a bit overwhelming. The good news: Your Chem I class doesn't have to



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5.24a KEY I. For each  
of the following  
statements, write “I” for  
ionic, “C” for covalent

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## Chemistry Unit 5

December 11, 2011 31.

Balance the following equations by putting the proper coefficients

wherever they are

needed. a)  $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$

b)  $\text{Al} + \text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_3$

+  $\text{H}_2$  c)  $\text{Al} + \text{O}_2$

$\rightarrow \text{Al}_2\text{O}_3$  d)  $\text{C}_5\text{H}_{12} + \text{O}_2$

$\rightarrow \text{CO}_2 + \text{H}_2\text{O}$  e)  $\text{Na} + \text{H}_2\text{O}$

$\rightarrow \text{H}_2 + \text{NaOH}$  f)

$\text{NaClO}_3 \rightarrow \text{NaCl} + \text{O}_2$  g)

...

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