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Chemistry, class-12, unit-7
~~Chemistry Review Sheet Unit 7~~

are representative of each type of

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problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes. 1. Calculate the molar mass for ethane, C₂H₆. C₂H₆ 2(12.0 g) + 6(1.0 g) = 30.0 g 2. Convert 45 g Se to moles of Se. ? mol Se = 45 g Se ! 1 mol Se 79.0 g Se = 0.57 mol Se 3.

~~Review Sheet: Unit 7 Name~~

~~Vocabulary: Fill in the blanks ...~~

are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes. 1. Calculate the molar mass for ethane, C₂H₆. 2. Convert 45 g Se to moles of Se. 3. Convert 4.77×10^{24} molecules of SO₂ to grams.

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~~Review Sheet: Unit 7 Name~~

~~Vocabulary: Fill in the blanks ...~~

UNIT 7 REVIEW SHEET –

EQUILIBRIUM • EQUILIBRIUM is the point at which the rate of the forward reaction is equal to the rate of the reverse reaction in a reversible process. o Reaction is dynamic – still occurring – even though to the eye, it looks like nothing is changing at equilibrium. o Observable/measurable quantities (concentration, pressure, color) are constant at equilibrium. o The amounts of products and reactants are typically not equal at equilibrium!

~~UNIT 7 REVIEW SHEET.pdf – UNIT 7~~

~~REVIEW SHEET \u2013~~

CHEMISTRY CP Name: _____ Per:

_____ TEST: _____ Unit 7 Review

Sheet: Thermochemistry KEY

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~~Answer~~ Potential Energy Diagrams: Choose the appropriate letter. 1. What is the ΔH of the forward reaction? f 2. What is the activation energy of the forward reaction? b 3. What is the ΔH of the reverse reaction? f 4.

~~Chemistry Unit 7 Review Sheet KEY
OAK PARK USD~~

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Review Sheet: Unit 5 Name _____

CHEMISTRY: A Study of Matter © 2004, GPB 5.24a KEY I. For each of the following statements, write "I" for ionic, "C" for covalent

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Answers

~~CHEMISTRY: A Study of Matter~~
Review Sheet for Chemistry Unit 5
December 11, 2011 31. Balance the following equations by putting the proper coefficients wherever they are needed. a) $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$ b) $\text{Al} + \text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_3 + \text{H}_2$ c) $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ d) $\text{C}_5\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ e) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{NaOH}$ f) $\text{NaClO}_3 \rightarrow \text{NaCl} + \text{O}_2$ g ...

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