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(inert for chemical reaction) is greatly simplified material balance calculations. This substance is termed as (tie component). It is important to search for the existence of a tie component and formulate a material balance equation. Nitrogen is considered as a tie component in

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a given process, 100 kmol of carbon is burned in a furnace. It has been found that 20% of the carbon undergoes incomplete combustion resulting to CO production.

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equation is applied to each chemical compound appearing in the process.

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system

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$2 + 3H_2 + 2NH_3 \rightarrow 3(1) H_2O = 92 \text{ kJ mol}^{-1}$
 $S^0 = 192 \text{ J mol}^{-1} \text{ K}^{-1}$
To find the Gibbs free
energy of formation at room

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temperature, recall that $G_0 = H_0 - TS_0$
 $S_0(2) = 92 \text{ kJ mol}^{-1} + (298 \text{ K}) 0.192 \text{ kJ mol}^{-1} \text{ K}^{-1} = 35 \text{ kJ mol}^{-1}$
Alternatively, one can also find the temperature for which $G = 0$, $T = H_0 / S_0 = 479 \text{ K} = 206 \text{ C}$.

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ChE471: Chemical Reaction
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Introduction to Engineering

Calculations Date: _____ The goal of
Chapter 2 is to introduce the student
to the basics needed to perform
engineering calculations, including

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units, unit conversions, conversion
between force and weight, numerical
accuracy and precision, dimensional
homogeneity, and data analysis.

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