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## **Nail and Copper II chloride experiment**

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Iron and Copper (II) Chloride Lab  
Notebook

### **Reaction Between Iron and Copper (II) Chloride**

~~Iron and Copper(II)Chloride  
procedure CuCl<sub>2</sub> and Iron Nail Lab  
Part 2~~

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Reaction of iron nails with copper sulphate solution

*Copper Sulfate and Iron Nail Lab: Part 1*

~~Timelapse of Iron Nail in~~

~~Copper(II) Chloride~~ Reaction of Iron Nails with Copper Sulphate Solution in Water - MeitY OLabs

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Chemistry Revision - Iron \u0026amp;

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Copper Sulphate solution *The Nail Lab - Introduction to Chemical Reactions* Chemical Reactions: The Nail Lab **Nickel Plating at Home - Easy Electrolysis** **Electroplating** *conditions necessary for rusting of iron experiment* **How to grow**

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**beautiful crystals of salt - do  
your chemical experiment!**

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Thermal Decomposition of  
Copper(II) Sulfate Pentahydrate  
~~Chemistry Experiment Episode 1~~  
~~Electrolysis of Copper Chloride~~

**Turn Pennies Silver and Gold  
(Chemistry Trick) Preparation**

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## **of Copper I and Copper II**

**Chloride** Aluminum + copper II  
chloride ~~Electrolytic Refining of  
Metals | #aumsum #kids~~

~~#science #education #children~~  
Aluminum and Copper (II)

Chloride Reaction Iron and  
Copper II Chloride Reaction

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**Copper Chloride and  
Aluminum Foil** Lab Demo -  
Copper (II) sulfate solution reacts  
with iron nail The Aluminum and  
Copper Chloride Reaction prelab  
~~Nail Lab Part 3~~ **Iron reacting  
with copper sulphate  
solution(Activity 3.12- NCERT**

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## - X Science - Karnataka)

Copper Chloride and Iron Moles of  
Iron and Copper Lab **Iron Nail**

### **Copper Chloride Lab**

An iron nail is placed in a copper (II) chloride solution. The Iron is massed before and after and the other product is massed after

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Answers. Initial mass of ...

## **Nail and Copper II chloride experiment - YouTube**

Conclusion. When 2 solid iron  
nails is added to a solution of  
copper (II) chloride, a  
displacement reaction will occur.

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Also, products of reaction are iron (III) chloride and copper (II). In this experiment the actual oxidation number of iron is 3. The percent yield of copper is approximately 164.2%.

## **Lab Report of Decomposition**

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## **Answers** **Reaction Between Iron and ...**

Be careful not to lose any of the solid. \* Step I: 1) Allow the copper in the beaker to dry at low temperature in a drying oven for at least one hour. 2) Determine the mass of the beaker and the copper. 3) Repeat the drying

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process just to be sure that the copper is completely dry, and again determine the mass of the copper and the beaker.

## **The Reaction of Iron Nails with a Copper Solution Essay**

...

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$\text{Fe (s)} + \text{CuCl}_2 \text{ (aq)} \rightarrow \text{FeCl}_2 \text{ (aq)} + \text{Cu (s)}$  This reaction has a 1:1 mole ratio between iron used and copper produced. The iron will be provided by two iron nails. The copper will initially be in the form of  $\text{Cu}^{2+}$  ions in solution. (That's what the "aq" means

# Read Online Iron Nail Copper Chloride Lab (above.)

## **Lab: Iron and Copper Exchange**

We dropped iron nails in copper chloride with water solution.

1.what are any precautions you need to take? 2.explain what and

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aqueous solution is and give an example from the lab? 3. when you decanted identify the solution and the precipitate? 4. why did you rinse the precipitate with water then hydrochloric acid and again with water? 5. please write the balanced chemical equation

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## **Iron Nails Copper Chloride Chemistry Lab Questions ...**

Iron comes before copper in the reactivity series. So there is no doubt that Iron will displace copper from copper chloride to

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Answers  
make Ferrous Chloride ( $\text{FeCl}_2$ ).  
Copper will be deposited on the  
iron...

## **Iron Nails and Copper Chloride Reaction.. Is it Fe 2+ or 3 ...**

Iron nails kept in copper sulphate

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Answers.  $\text{Fe (s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{Cu (s)} + \text{FeSO}_4(\text{aq})$  When an iron nail is dipped in copper sulphate solution, a brown coating of copper is formed on the surface of iron and the colour of copper sulphate solution changes from blue to light green. This reaction

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Answers  
Shows that iron is more reactive than copper as it displaces copper from its solution and iron passes into solution as  $\text{Fe}^{2+}$  ions and ferrous sulphate solution is formed.

**Iron nails kept in copper**

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## **Answers sulphate solution « Science Labs**

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Answers  
iii) Add the two iron nails to the Erlenmeyer flask containing approximately 50ml of copper (II) chloride solution. iv) Place a rubber stopper on the Erlenmeyer flask to stop the solution from evaporating. v) Leave the reaction for at least one night.

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## **Decomposition Reaction Between Iron And Copper (II) Chloride**

Formal Lab: Iron, Copper, and  
Stoichiometry This lab will be an  
attempt to get the highest  
possible percent yield in

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performing a single replacement reaction. You'll be taking an iron nail and placing it in a copper (I) chloride solution. The result will be pure copper metal.

## **Formal Lab: Iron, Copper, and Stoichiometry**

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**Answers:** iron + copper sulphate  
= copper + iron sulphate Under  
the conditions of the reaction,  
there would be very small  
amounts of iron(III) sulphate but  
the main product would be the  
iron(II)...

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## **What happen when you put Sodium Chloride, Blue Copper**

...

This reaction has a 1:1 mole ratio between iron used and copper produced. The iron will be provided by two iron nails. The copper will initially be in the form

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of  $\text{Cu}^{2+}$  ions in solution. (That's what the "aq" means above.) 2. What does it mean to be aqueous? The iron nails and an empty beaker will be weighed before the reaction takes place. The nails will be cleaned of copper after the reaction ceases,

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dried and then re-weighed. We expect their weight to go down.

## **Moles of Iron and Copper - d3 2og0qmya1dw8.cloudfront.net**

An iron nail was dipped in a salt solution. After sometime a

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Answers  
reddish brown deposition of the nail was seen. The salt solution could be (a) Silver nitrate (b) Sodium sulphate (c) Aluminium chloride (d) Copper sulphate. 24. The colour of the residue left in the test tube after heating ferrous sulphate which undergoes

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decomposition is: (a) yellowish-brown

## **NCERT Class 10 Science Lab Manual Types of Reactions ...**

Hypothesis: It was hypothesised that if the iron nails were placed in higher concentrations of salt

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Answers  
water then the nail in the strongest concentration of sodium chloride would have the highest rate of corrosion and thus corroding more quickly than the nails in lower concentrations. This is because the higher concentrations of sodium chloride

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Answers contain more ions ( $\text{Na}^+$   $\text{Cl}^-$ ) that assist in ...

## **Effect of Sodium Chloride (NaCl) on Rust: Lab Explained**

...

When an iron nail is dipped in copper sulphate solution, a brown

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Answers  
A coating of copper is formed on the surface of iron and the colour of copper sulphate solution changes from blue to light green.

Core Science Lab Manual with

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