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Finding Molarity,

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Limiting Reactant

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Stoichiometry
Calculation Practice
Worksheet. 1.

Calculate the number
of moles of NaOH that
are needed to react
with 500.0 g of H. 2.
SO. 4. according to

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Calculations
the following
equation: H. 2.

Worksheet

Answers

Stoichiometry Calculation Practice Worksheet

Answer the following stoichiometry-related questions: 12) Write the balanced equation for the reaction of acetic acid with aluminum hydroxide to form water and

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aluminum acetate: 13)

Using the equation
from problem #12,
determine the mass of
aluminum acetate that
can be made if I do
this reaction with 125
grams of acetic acid

Stoichiometry Practice Worksheet

Stoichiometric

Calculations 1.

Sodium metal burns

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Calculation
Worksheet
Answers

in air according to the balanced reaction shown below. $4 \text{Na} + \text{O}_2(\text{g}) \rightarrow 2 \text{Na}_2\text{O}$ Complete the setups with the correct factors to answer the following questions: (a) How many moles of oxygen are needed to completely react with 9.5 g of sodium? mol O_2 (b) How many grams of

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Calculations are needed to produce 12.5 g of sodium oxide? 12.5 g Na₂O x x 62.0 g 2.

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Stoichiometric Gram
to Gram Calculations
Worksheet - Answers.

1. $2C_4H_{10} + 13 O_2$
----> $8 CO_2 + 10$
 H_2O . 1. (a) Find the
moles of water that

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were formed. $n = m =$
 $2.46\text{g} = 0.14$ moles of
water formed. M

18.02 g/mol . 1. (b)

From the balanced
equation the reaction
ratio is.

Stoichiometric Worksheet #2: Gram to Gram Calculations

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Calculations.

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Calculations -

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Worksheet for Basic
Stoichiometry. Part 1:
Mole ?? Mass

Answers. Convert
the following number
of moles of chemical
into its corresponding
mass in grams. 1.

0.436 moles of
ammonium chloride.

2. 2.360 moles of lead
(II) oxide. 3. 0.031
moles of aluminum
iodide.

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Worksheet for Basic Stoichiometry

CHM 130

Stoichiometry

Worksheet. The

following flow chart

may help you work

stoichiometry

problems. Remember

to pay careful

attention to what you

are given, and what

you are trying to find.

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1. Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide:
C.

CHM 130 Stoichiometry Worksheet

2. , would be. Ca:
1(40.1 amu) + Cl:
2(35.5 amu) 111.1

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amu. •Formula

weights are generally reported for ionic compounds.

Stoichiometry. ©

2009, Prentice-Hall, Inc. Molecular Weight

(MW) •A molecular weight is the sum of the atomic weights of the atoms in a molecule. •For the molecule ethane, C.

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Stoichiometry: Calculations with Chemical Formulas and ...

To solve stoichiometry problems with limiting reactant or limiting reagent: 1. Figure out which of the reactants is the limiting reactant or limiting reagent. 2. See how much product can be

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formed by using the maximum amount of the limiting reactant or limiting reagent. 3.

Stoichiometry - Limiting and Excess Reactant (solutions

...

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According to the balanced chemical equation, 6 mol of CO₂ is produced per mole of glucose; the mole ratio of CO₂ to glucose is therefore 6:1. The number of moles of CO₂ produced is thus.

$$\begin{aligned} & (5.3.3) \text{ moles CO}_2 \\ & = \text{mol glucose} \times \\ & 6 \text{ mol CO}_2 / 1 \text{ mol} \end{aligned}$$

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5.3: Stoichiometry Calculations -

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25 mol \times 3 = 75 moles
of oxygen. The
burning of 25 moles of
ethyl alcohol requires
75 moles of oxygen.

3. (b) From the
balanced equation the
reaction ratio is: 1

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mol x = 10 moles of
alcohol. The

consumption of 30
moles of oxygen
requires 10 moles of
ethyl alcohol.

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About This Quiz &
Worksheet The
questions will mainly

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deal with definitions of key terms. These questions will give you a definition and you will need to select the correct term. Other questions...

Quiz & Worksheet - Mass-to-Mass Stoichiometric ...

Solutions for the
Stoichiometry

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1) 355.3 grams of

Na_2SO_4 2) 313.6

grams of LiNO_3

3) a) How many liters of 0.100 M HCl would be required to react completely with 5.00 grams of calcium hydroxide?

**Stoichiometry
Practice Worksheet**
MOLES MOLES. $x\text{A}$
 $y\text{B} + z\text{C}$. GIVEN:

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WANTED: Grams A x
1 mole A x y mole B x
g B = Gram B. g A x
mole A 1 mole B.

molar mass A mole
ratio from molar mass
B. the balanced
equation. Double
lined boxes are
Conversion Factors to
convert from one
quantity to another.
mole.

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